Summary of Bonds and Forces

	Ionic Bond	Covalent bond	Hydrogen Bond	Dipole-dipole force	London forces
Where do you see it?	Between non-metal and metal in an ionic compound	Between non-metal and non-metal in a covalent compound	Special type of dipole/dipole force found between two polar molecules possessing OH, NH or FH bond	Occurs between two polar molecules	Occurs between two atoms or molecules due to temporarily migration of e-, causing temporarily dipole → Always present between two molecules
Strength	Very strong	Very strong	Stronger than dip/dip	As strong as London	As strong as dip/dip
Example	KCl	CH_4	H_2O and H_2O	CH_3F and CH_3F	H_2 and H_2